

Name \_\_\_\_\_

## Noble Gases: The Mysterious and Chemically Unreactive Elements

### Multiple Choice Questions

1. What is the main reason noble gases are chemically unreactive?
  - a) They have high electronegativity.
  - b) They have complete electron configurations.
  - c) They readily gain or lose electrons.
  - d) They are highly reactive metals.
  
2. Which of the following elements is not a noble gas?
  - a) Helium
  - b) Neon
  - c) Argon
  - d) Oxygen
  
3. What is the octet rule?
  - a) The tendency of elements to lose electrons in chemical reactions.
  - b) The tendency of elements to gain electrons in chemical reactions.
  - c) The rule that elements should have eight electrons in their outermost energy level to be stable.
  - d) The rule that elements should have only two electrons in their outermost energy level to be stable.
  
4. Why does helium differ from other noble gases in terms of reactivity?
  - a) It has a smaller atomic size.
  - b) It lacks a full complement of electrons in its outermost energy level.
  - c) It readily forms stable bonds with other elements.
  - d) It is not a noble gas.
  
5. Which noble gas is commonly used in neon signs and lighting?
  - a) Argon
  - b) Radon
  - c) Krypton
  - d) Neon

