

Name _____

How Do Redox (oxidation-reduction) Reactions Work?

Multiple Choice Questions

1. What happens during oxidation in a redox reaction?
 - a) A substance gains electrons.
 - b) A substance loses electrons.
 - c) Oxygen is added to a substance.
 - d) Hydrogen is removed from a substance.

2. What is reduction in a redox reaction?
 - a) A substance gains electrons.
 - b) A substance loses electrons.
 - c) Oxygen is added to a substance.
 - d) Hydrogen is removed from a substance.

3. In a redox reaction, what is the simultaneous occurrence of oxidation and reduction in two substances called?
 - a) Fusion
 - b) Combustion
 - c) Double replacement
 - d) Redox reaction

4. Which of the following is an application of redox reactions?
 - a) Combustion
 - b) Boiling water
 - c) Solubility
 - d) Refraction

5. How do redox reactions contribute to rusting?
 - a) By preventing the reaction between iron and oxygen
 - b) By preventing the reaction between iron and water
 - c) By facilitating the oxidation of iron
 - d) By neutralizing oxygen molecules

