Name _____

Entropy Unveiled: Exploring the Second Law of Thermodynamics

Multiple Choice Questions

1. What does the concept of entropy measure?

a) The amount of usable energy in a system

b) The level of order or disorder in a system

c) The temperature of a system

d) The number of particles in a system

2. According to the second law of thermodynamics, what happens to entropy in a closed system over time?

a) Entropy always decreases.

b) Entropy remains constant.

c) Entropy increases.

d) Entropy fluctuates randomly.

3. How can you observe the concept of entropy in your room?

a) Your room becomes messier over time.

b) Your room becomes cleaner over time.

c) Your room stays exactly the same.

d) Your room gets rearranged daily.

4. What is the term for changes that can be reversed without increasing overall entropy?

a) Irreversible changes

b) Chaotic changes

c) Random changes

d) Reversible changes

5. Which of the following is an example of an irreversible change?

a) Melting ice

b) Shuffling a deck of cards

c) Spreading perfume scent

d) Closing a door

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